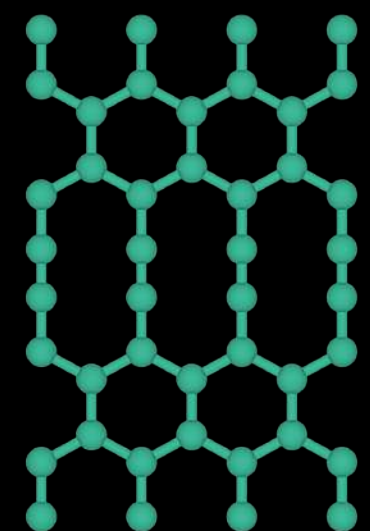


On the CO₂ Harvesting from N₂ Using Grazyne Membranes

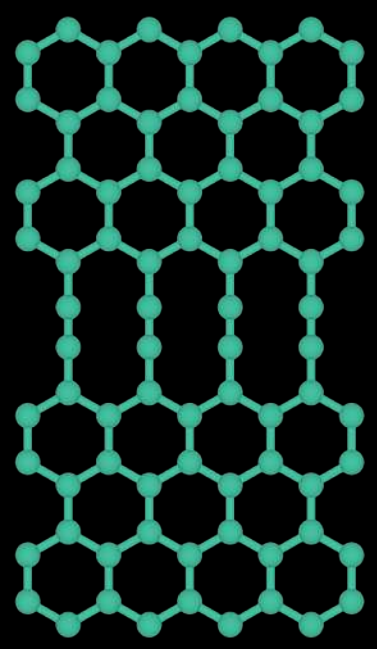
WE PROPOSE GRAZYNES

Grazynes are a subtype of graphynes, 2D single-layer C-based materials with sp and sp² C atoms hybridizations. In grazynes, graphene-like stripes with sp² hybridized C atoms are interconnected by acetylenic linkages with sp C atoms.

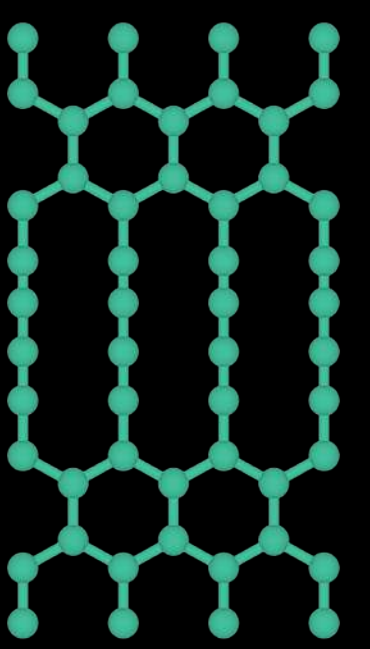


[1],[1]-grazyne

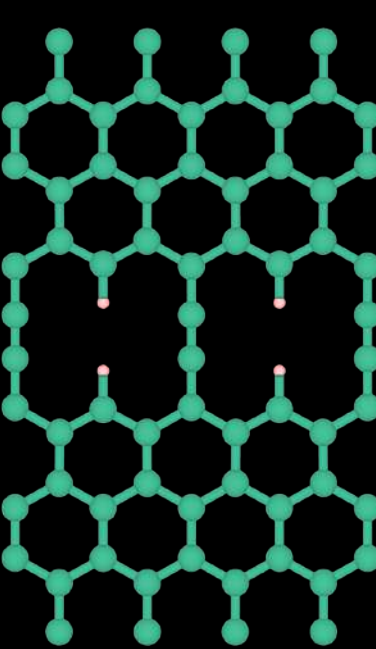
They can be modified in different ways



Modifying the graphene width

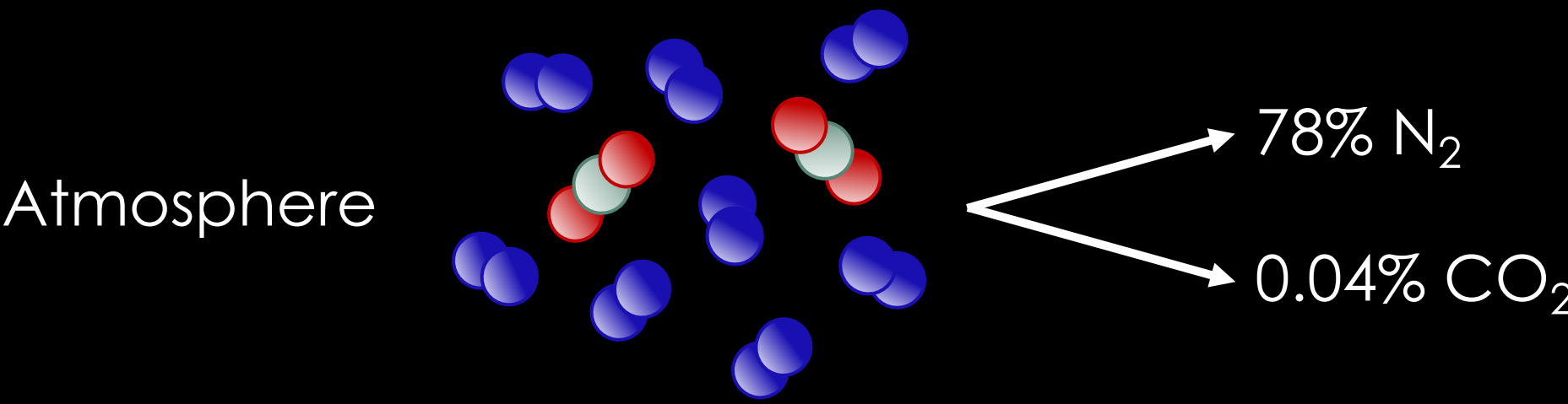


Enlarging acetylenic linkages



Creating acetylenic vacancies

The increase in the concentration of CO₂ in the atmosphere is one of the main drivers of climate change. In this context, the separation of CO₂ from N₂ is a critical process.

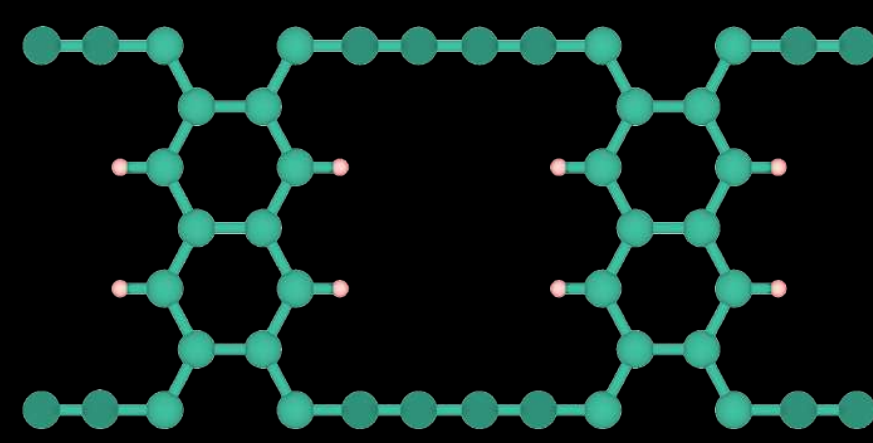


To effectively use this N₂, it must be in a high-purity state, which requires purification processes. Within the existing options, membrane-based separation stands out as a promising solution for CO₂ separation, boosting numerous benefits such as environmental friendliness, a high active surface area, and easy maintenance.

Air Cleaner

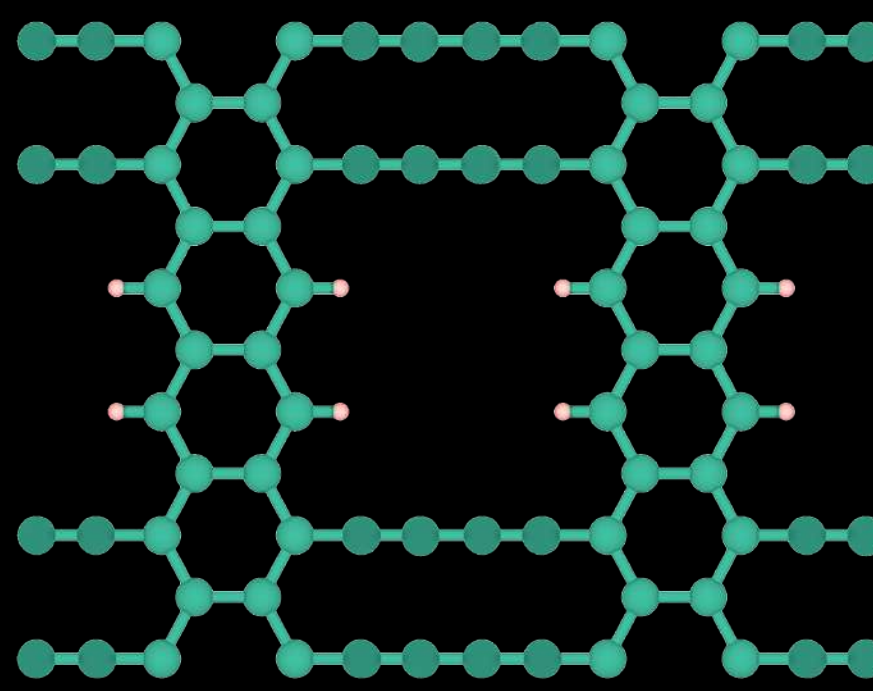


GRAZYNE MODELS



[1],[2]{2} ⇌ A

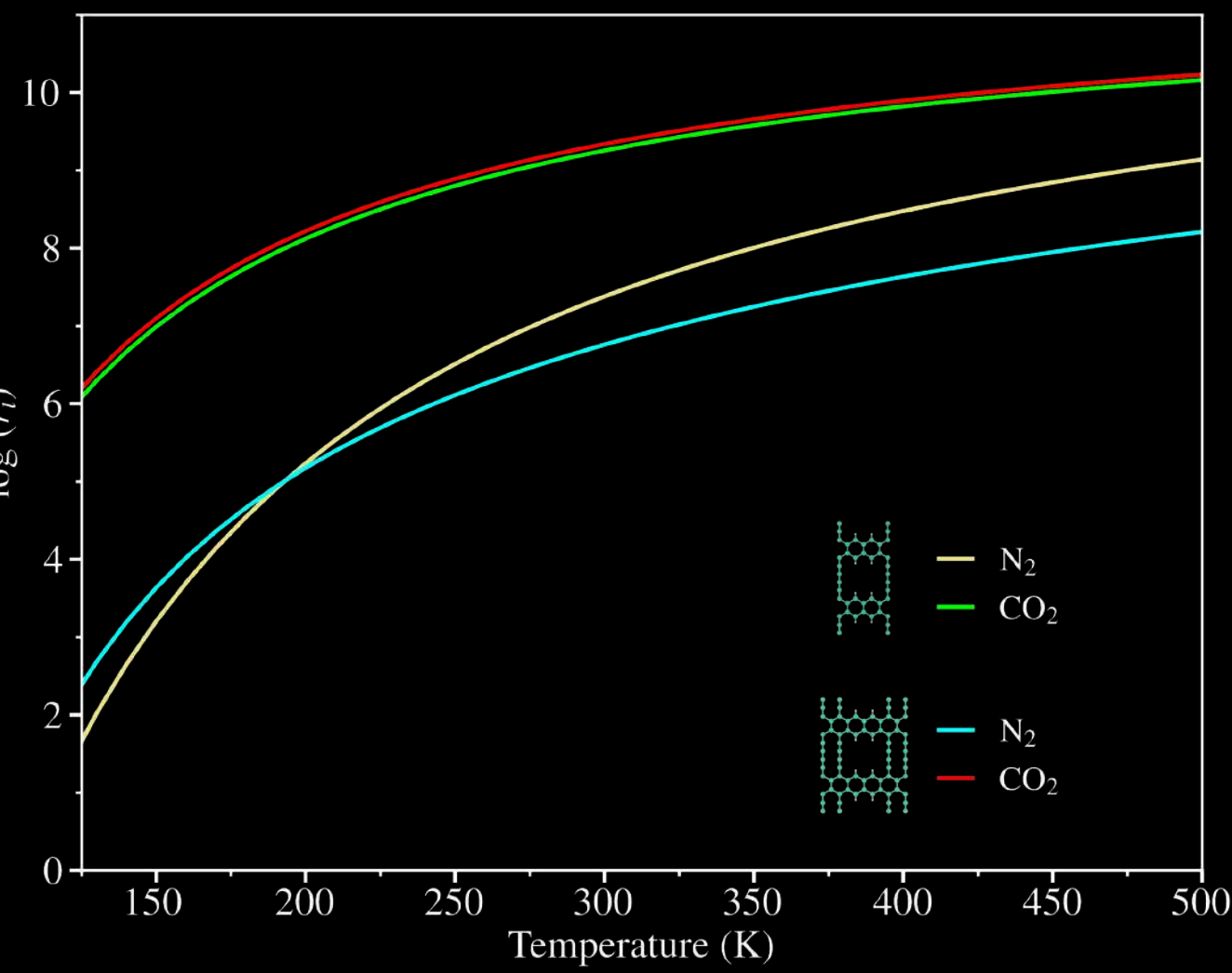
Composed of one-ring wide graphene stripes interconnected by acetylenic bonds, spanning a distance equivalent to two triple bonds, with two acetylenic vacancies.



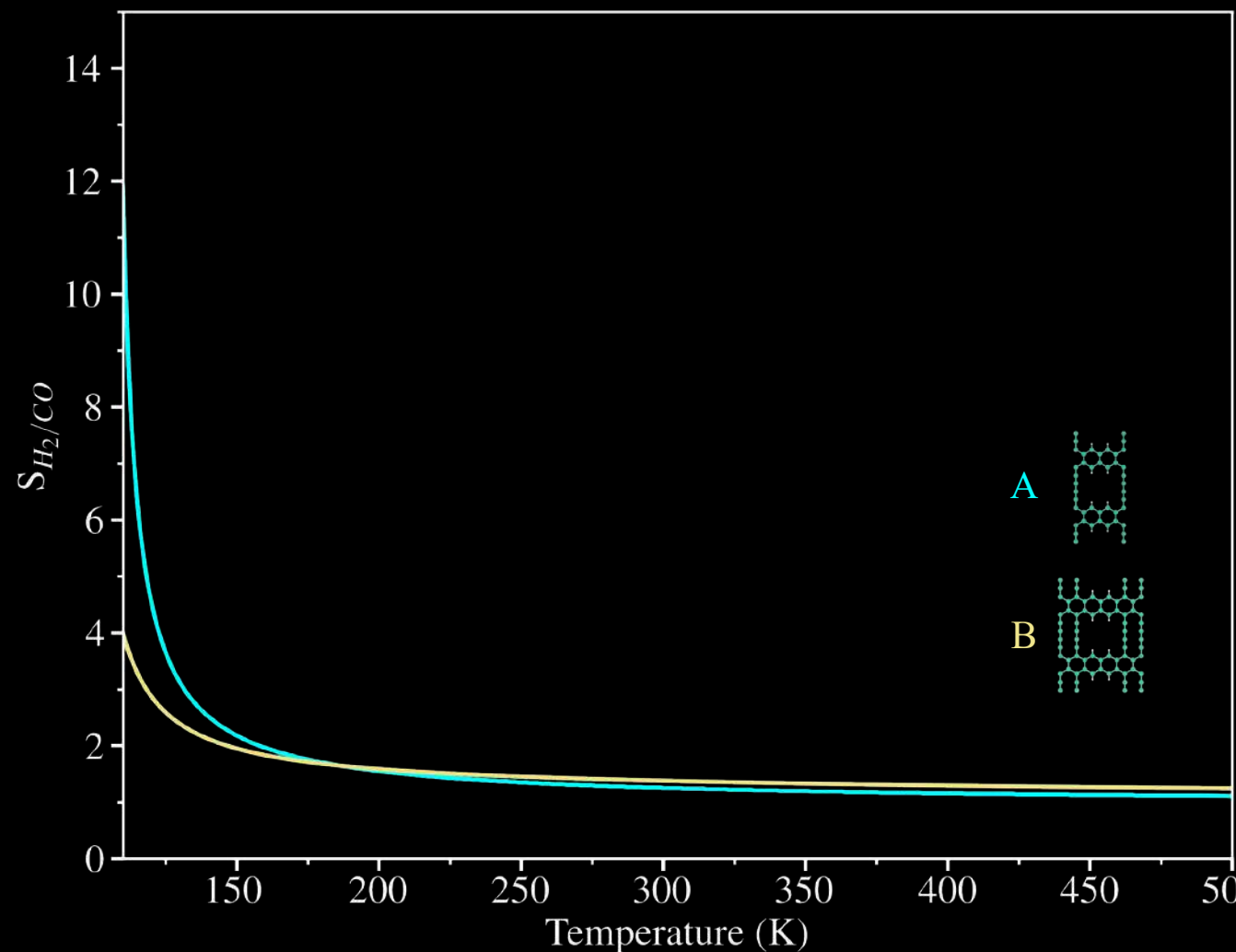
[1],[2]{(0,0),2} ⇌ B

The difference between the structures lies in the A-grazyne containing a solitary acetylenic bond between consecutive pores, while the B-grazyne exhibits two adjacent acetylenic bonds followed by two vacancies.

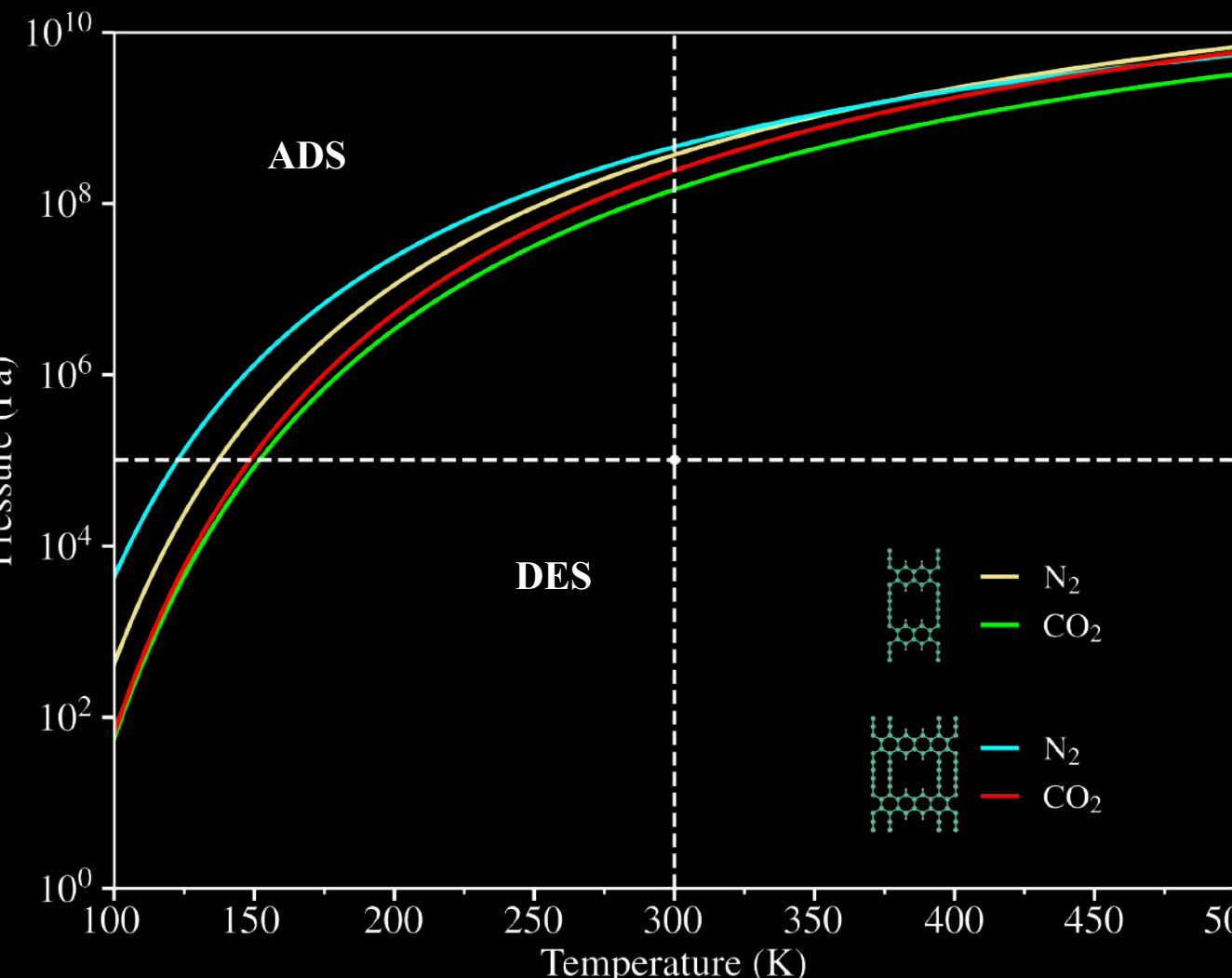
DFT RESULTS



CO₂ consistently shows higher membrane diffusion rates than N₂, regardless of the crazyne model. The figure highlights a key temperature range between 100 and 500 K, where diffusion rates, r_i , exceed 1 s⁻¹, suggesting the potential use of grazynes as molecular separation membranes.



The analysis of the selectivity confirms that CO₂ has a higher selectivity than N₂, but this decreases as temperature goes up, following the rate constants, with a clear trend toward unity at higher temperatures. Therefore, lower temperatures are better for separating CO₂ from N₂, though this comes with a lower permeation rate.



The KPD highlights two distinct regions: one where adsorption is favored, and another where desorption is preferred. CO₂ shows a wider adsorption range. At p=1 atm and T=300 K, neither CO₂ nor N₂ is inclined toward adsorption, making grazynes effective for separation membranes.

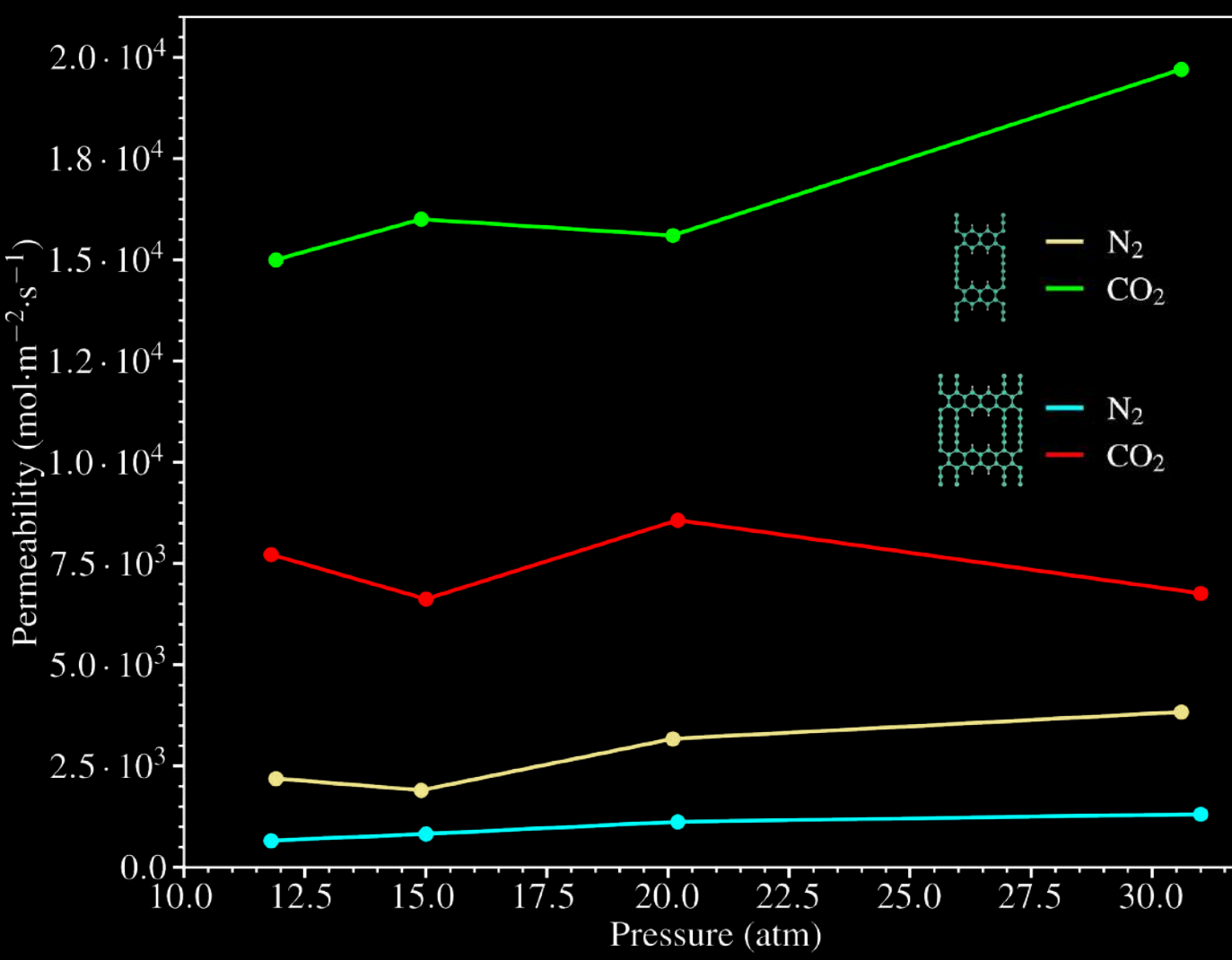
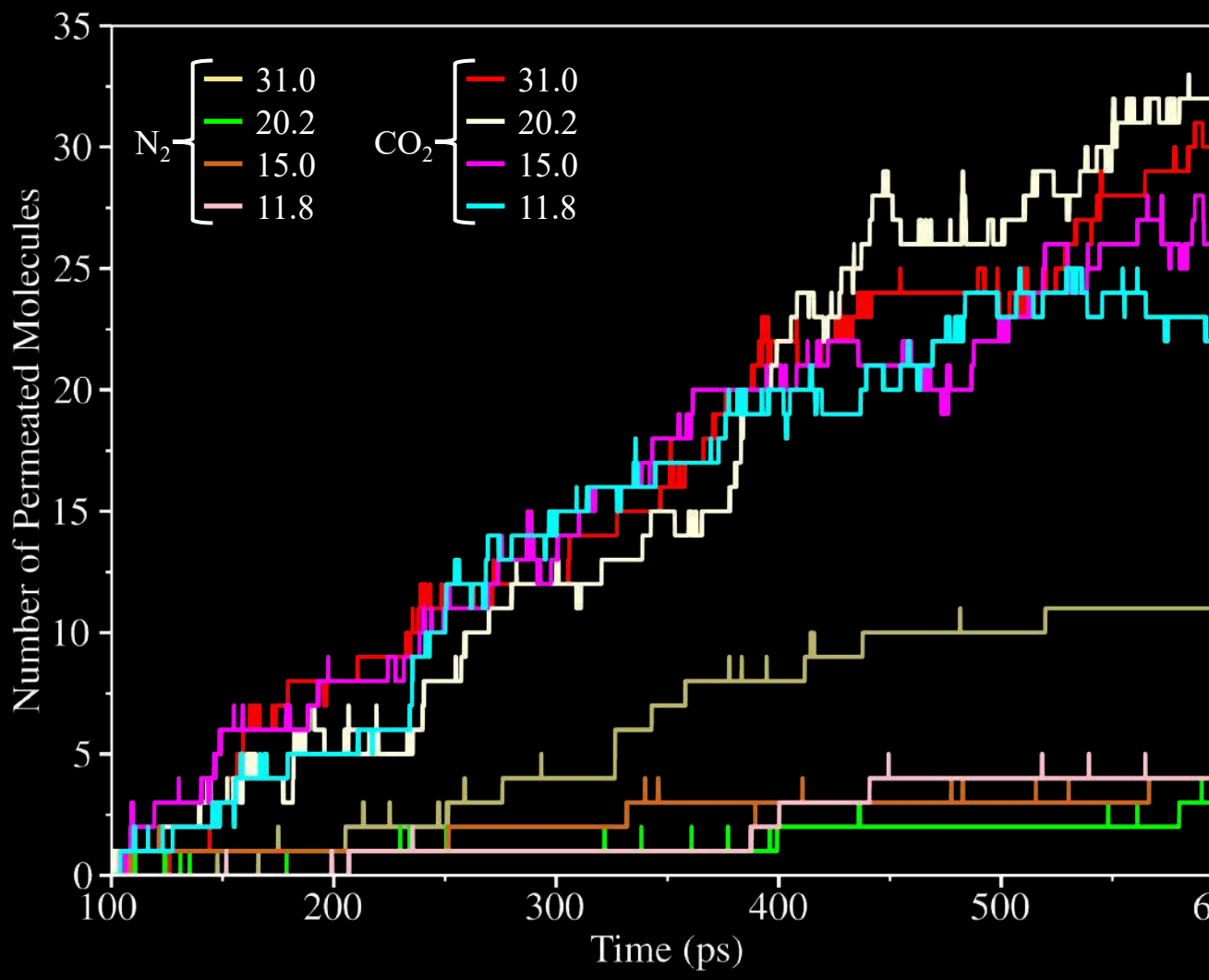
CONCLUSIONS

DFT and MD support the use of crazyne membranes for selective CO₂ separation from N₂. Grazynes are capable of physisorbing CO₂ and N₂, thus avoiding material poisoning by molecular decoration, while the diffusion of CO₂ through the pores is found to be rapid, yet easier than that of N₂, in the rate order of the s⁻¹ in the 100-500 K temperature range. In addition, according to MD with equimolecular mixtures of CO₂:N₂, ca. 60-70% of CO₂ permeates the A-grazyne, with the remaining 30-40% being N₂. Conversely, B-grazyne achieves up to 90% CO₂ permeation.

MOLECULAR DYNAMICS RESULTS

Grazyne	A	B
Pressure (atm)	30.6	20.1
500 ps	18	14
1500 ps	31	31
#N ₂	39	41
#CO ₂	30	30
#N ₂	23	9
#CO ₂	45	38
#N ₂	4	4
#CO ₂	22	13
#N ₂	13	13
#CO ₂	33	33

The number of CO₂ (#CO₂) and N₂ (#N₂) molecules that passed through shows that CO₂ diffuses more easily through both crazyne membranes. CO₂ makes up about 65-75% of the total molecules permeating through A-grazyne and 70-90% for B-grazyne. These results suggest a potential CO₂ enrichment of ca. 90%. This trend is also seen in the time evolution of the number of molecules that passed through B-grazyne, while a similar trend is observed for A-grazyne (not shown). Generally, increasing the pressure leads to more permeation, though some trends may fluctuate. It is important to remember that the likelihood of a molecule passing through a pore depends on its arrival orientation. No parallel permeation has been observed, thus molecules approaching perpendicularly have a higher chance of crossing.



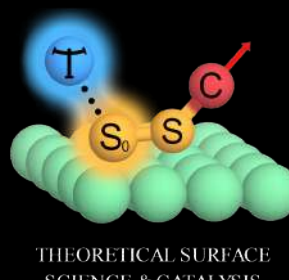
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S. Kamalinahad, F. Viñes, P. Gamallo, *J. Phys. Chem. C* **2019**, 123, 27140–27149.
F. Viñes, A. Calzada, P. Gamallo, *J. CO₂ Util.* **2023**, 71, 102459.

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